

Review Unit 4-Chapter 11.3, 15.1,15.2, Ch 7, 16.1 and 16.2 Chem 11- Mrs. Sanford Name:						
Spectator ion-						
Net ionic equation-						
Using a solubility table salt of alkali metals are_____, ammonia is also soluble. Nitrates, _____, sulphates and chlorides(except for those attached to Pb^{2+} , Ag^+ , Ba^{2+} , Sr^{2+} , Ca^{2+} . Insoluble compounds usually contain _____, _____, _____, _____and hydroxides. (table 11.3)						
Water is polar because of _____.						
There are_____valence orbitals in the valence level of representative elements.						
Surface tension-						
Vapor pressure-						
Solute-						
Solvent-						
Solvation-						
Electrolyte-			Non-electrolyte-			
Hydrate-						
#valence electrons=_____						
A maximum of _____electrons can occupy each individual orbital in the valence orbitals.						
Cation-			Anion-			
Octet rule:						
Lone pair: (or non-bonding pair)			Bonding pair:			
Draw Lewis electron dot diagrams below for the following elements: H, Mg, I, S, P, C, B						
H•						
Ionic bond-						
Chemical formula-						
Formula Unit-						
Three properties of ionic compounds are-						
1-		2-		3-		
Metallic bonds-			An alloy is-			

1- _____ 2- _____ 3- _____

[illegible][illegible][illegible]

